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| 1.  pH = -Log [H+] | a.  A substance where 100% of the acid is ionised (dissociated) |
| 2.  Strong acid | b.  The acid dissociation constant (symbol for) – used to work out how much acid is dissociated |
| 3.  H30+ | c.  The general formula for an acid |
| 4.  HA | d. The equation used to work out the pH of a substance. |
| 5.  Ka | e.  hydroxium ion |
| 6.  Ka = [H+][A-]  [HA] | f.  Symbol to show that the reaction is reversible. |
| 7. | g.  An acid where not all of the acid is ionised (dissociated) |
| 8.  Weak acid | h.  Ka expression (equation) – the quantitative measure of the strength of an acid in solution |