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| 1.pH = -Log [H+] | a.A substance where 100% of the acid is ionised (dissociated) |
| 2.Strong acid | b.The acid dissociation constant (symbol for) – used to work out how much acid is dissociated |
| 3.H30+ | c.The general formula for an acid |
| 4.HA | d. The equation used to work out the pH of a substance. |
| 5.Ka | e.hydroxium ion |
| 6.Ka = [H+][A-] [HA] | f.Symbol to show that the reaction is reversible. |
| 7. | g.An acid where not all of the acid is ionised (dissociated) |
| 8.Weak acid | h.Ka expression (equation) – the quantitative measure of the strength of an acid in solution |